PtBamnCe/Al as a High-Efficiency Catalyst for Nox Storage and Reduction as a Function of Cycling Conditions

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Abstract

The aim of the study was to present novel PtBamnCe/Al material as the high-efficiency catalyst for NOx storage and reduction (NSR) application. In the first part of this paper, the performance of PtBamnCe/Al catalyst on the NOx conversion to N2 was presented. Compared with traditional model PtBa/Al catalyst, the PtBamnCe/Al catalyst exhibit at least 2 times higher in NOx conversion efficiency at 400 °C. More interesting, ammonia yield on PtBamnCe/Al is limited to a very lower level (nil). With 6% H2 in the rich pulse, the 94% of NOx conversion to N2 can be obtained with PtBamnCe/Al catalyst at 400 °C. The second part deals with sulfur resistance of PtBamnCe/Al catalyst. Sulfur poisoning, regeneration of sulfated catalysts and NOx removal efficiency under rich/lean cycling conditions are studied. PtBamnCe/Al is more sulfur resistant than PtBa/Al. However, the sulfur poisoning is irreversible and only 85% of the initial NOx conversion could be recovered after the regeneration treatment.

Keywords: Emission Control; Lean Burn; NOx Storage and Reduction; Sulfur in Fuel; Lean/-rich cycle

Introduction

Lower CO2 emissions from automotive sources are necessary and lead to the development of diesel and lean-burn engines. However, exhaust gases from these engines contain NOx in excess of O2, which makes NOx reduction into N2 very difficult. One possible solution is the use of the NOx storage reduction (NSR) catalyst [-3], working in transient periods: during the lean condition, NOx are firstly oxidized and stored as nitrites or nitrates on a basic material, usually barium oxide. Periodically, the catalyst is regenerated: the stored NOx(NO2) are reduced in N2 during a short excursion in rich condition. Nevertheless, the major drawback of this system is the deactivation of the catalyst, mainly due to sulfur poisoning [2,3], and ammonia emission can be formed during the short excursion under rich conditions, especially when hydrogen is used as reducer [-6].

In our previous study [5], ammonia intermediate pathway was clearly demonstrated for the reduction of the stored NOx with H2. It appears that, when hydrogen is missing in the rich pulses, that are fully converted, the ammonia selectivity tends to be nil because the produced NH3 can react with the remaining stored NOx. In opposition, if some hydrogen remains, the ammonia selectivity increases with the amount of excessive hydrogen. It induces that NOx reduction with H2 into ammonia is faster than the NOx selective catalytic reduction with ammonia. This is a fundamental result for our next step: catalyst modification for a better NOx conversion to N2.

In the first part of this paper, our latest result in catalyst improvement with a multi-component PtBamnCe/Al sample will be presented. In the second part of this paper, the impact of sulfur in gasoline on the performance of NSR catalyst will be discussed. Sulfur poisoning, regeneration of sulfated catalysts and NOx removal efficiency under rich/lean cycling conditions are studied [7].

Experimental

Catalyst preparation

The detailed preparation protocols are reported in our previous studies [8]. The reference catalyst contains 1 wt% Pt and 20 wt% BaO supported on alumina. Alumina powder (230 m2g-1) was immersed in an ammonia solution and was firstly impregnated using a barium nitrate salt. After evaporation at 80 °C and drying at 120 °C, the obtained powder was treated at 700°C under synthetic dry air. Platinum was then impregnated using a Pt (NH3)2(NO3)2 aqueous solution. After drying, the catalyst was pre-treated at 700 °C for 4h under N2 and finally
stabilized at 700 oC for 4h under a mixture containing 10% O2, 5% H2O in N2. The modified samples were prepared using the same protocol except that the nitrate salts of Mn(IV), and Ce (III) were simultaneously added with the barium salt. In this case, a part of alumina was replaced to assure the desired “additive/Ba” molar ratio. The catalysts containing both Mn and Ce were also prepared. In this case, Mn/Ba and Ce/Ba molar ratio are always 1. The Mn–Ce modified catalysts are noted PtBaMnCe/Al.

Catalytic activity measurements: NOx conversion in cycling conditions

<table>
<thead>
<tr>
<th>Gas</th>
<th>NO</th>
<th>H2</th>
<th>O2</th>
<th>CO2</th>
<th>H2O</th>
<th>N2</th>
</tr>
</thead>
<tbody>
<tr>
<td>Rich</td>
<td>-</td>
<td>3-6%</td>
<td>-</td>
<td>10%</td>
<td>10%</td>
<td>Balance</td>
</tr>
<tr>
<td>Lean</td>
<td>500 ppm</td>
<td>-</td>
<td>10%</td>
<td>10%</td>
<td>10%</td>
<td>Balance</td>
</tr>
</tbody>
</table>

Before measurement, the catalyst (60mg) was pre-treated in situ at 450 oC under 3% H2, 10% H2O, 10% CO2 and N2 for 15min. The sample was then cooled down to 400 oC under the same mixture. The NOx conversion was studied in cycling condition by alternatively switching between lean (60s) and rich (3s) conditions using electro-valves. The gas composition is described (Table 1). NO and NO2 were followed by chemiluminescence, N2O by mass spectrometry. Before the analyzers, H2O was trapped in a condenser at 0 oC. The trapped water was analyzed by two different HPLCs for NH4+, NO2− and NO3−. NO2− and NO3− were added to the unconverted NOx. The N2 selectivity is calculated assuming no other N-compounds than NO, NO2, N2O, NH3.

Results and Discussions

NOx storage-reduction efficiency of PtBaMnCeX/Al catalyst

The influence of cerium-manganese addition on the NOx storage-reduction efficiency of PtBa/Al was studied at 400 oC using first 3% H2 in the rich pulses. For comparison, results obtained with the PtBa/Al reference catalyst are also reported in (Figures 1 & 2). At 400 oC with 3% H2 in the rich pulses, the hydrogen conversion reaches 100% with PtBaMnCe/Al catalyst (Table 2). Hence, the NOx conversion could be limited by the reducer amount. However, the NOx conversion to NH3 over this sample is higher than PtBa/Al catalyst. More interestingly, the ammonia selectivity becomes nil with PtBaMnCe/Al catalyst.

Table 2: NOx storage/reduction efficiency test at 400 oC with 3% H2 (a) or 6% H2 (b) in the rich pulses.

<table>
<thead>
<tr>
<th>Catalyst</th>
<th>H2 conversion (%)</th>
<th>NOx conversion (%)</th>
<th>N2 selectivity (%)</th>
<th>NH3 selectivity (%)</th>
<th>H2 conversion</th>
<th>NOx conversion (%)</th>
<th>N2 selectivity (%)</th>
<th>NH3 selectivity (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>PtBa/Al</td>
<td>85</td>
<td>45</td>
<td>67</td>
<td>33</td>
<td>63</td>
<td>48</td>
<td>57</td>
<td>43</td>
</tr>
<tr>
<td>PtBaMnCe/Al</td>
<td>100</td>
<td>70</td>
<td>100</td>
<td>0</td>
<td>99</td>
<td>94</td>
<td>100</td>
<td>0</td>
</tr>
</tbody>
</table>

Thus, with this condition with full H2 conversion, the in situ formed ammonia is able to react with the remaining stored NOx to produce N2. With 6% H2 in the rich pulses (Table 1), a significant NOx conversion improvement is observed with PtBaMnCe/Al sample, from 70% to 94%. These values can be considered as optimal values since the NOx storage rates of this sample reach 96% (result not shown). For the reference catalyst, the increase of H2 content in the rich pulse only leads to an increase in NH3 selectivity. This study shows that PtBaMnCe/Al catalyst is very promising for NOx storage-reduction application. Most importantly, the NH3 yield is very low (nil) with this catalyst. Mn-Ce addition in NSR catalyst leads to the enhancement of the catalytic properties. It can be attributed to an improvement of the reaction between the ammonia formed in situ and the remaining stored NOx [6].

Figure 1: NOx storage/reduction efficiency of PtBa/Al and PtBaMnCe/Al catalysts, test at 400 °C with 6% H2 (b) in the rich pulses. NOx conversion (%) into N2 ( ) and into NH3 (■): impact of sulfur poisoning and catalyst regeneration.

Figure 2: TPR profile of PtBa/Al (++) stabilized catalysts, (–) sulfated catalysts; (—) regenerated catalysts.
Impact of sulfur in gasoline on the performance of model NSR catalyst

In this section, the impact of sulfur on the performance of model NSR catalyst is focused. Firstly, the NSR catalyst was sulfated before catalytic activity measurement in order to simulate the deposition of sulfur compound in gasoline onto the catalyst surface. And then, the regeneration of sulfated catalyst was studied. The sulfating treatment was performed at 300 °C with SO₂ and corresponds to a 2.0 wt% S theoretical content if all the sulfur is stored on the catalyst. The regeneration of sulfated catalysts was carried out at 550 °C for 30 min under rich conditions with a mixture containing 2.5% H₂, 10% CO₂, and 10% H₂O in N₂.

SO₂ poisoning and regeneration

The NOx conversion rate of the fresh, sulfated and after sulfur regeneration treatment of Pt20Ba/Al and Pt20BaMnCe1/Al catalysts, measured at 400 °C, are reported in Figure 1. After the sulfating treatment, the NOx storage-reduction capacities strongly decrease. For Pt20Ba/Al, a loss of 56% is observed. Mn-Ce doped sample is more resistant toward the poisoning treatment, with a loss of 41%. After the regeneration treatment, only 79% of the initial value is recovered over Pt20Ba/Al, indicating that an efficient regeneration of the NOx storage sites cannot be obtained. This rate is slightly higher for Pt20BaMnCe/Al (85%) [9,10].

The sulfated Pt20Ba/Al catalyst was characterized by H₂-TPR (Figure 2). The sample was first in situ pretreated at 300 °C under pure oxygen. The reduction was carried out with 1% H₂ from room temperature up to 800 °C. As described in [4], the TPR profile of the sulfated Pt/10Ba/Al catalyst exhibits two main peaks around 500 and 600 °C. The first one is attributed to the simultaneous reduction of aluminum sulfates and some well dispersed barium sulfates located in platinum proximity. The second peak, around 600 °C, is ascribed to the reduction of surface barium sulfates, and the observed shoulder near 750 °C corresponds to bulk barium sulfate reduction.

Assuming a H₂/SO₂ 4/5 ratio of 4 for the sulfate reduction (XSO₄ + 4H₂ --- X-S + 4H₂O and/or XSO₃ + 4H₂ --- X-O + H₂S + 3H₂O), total sulfur content can be deduced from hydrogen consumption. After the sulfur regeneration, only 58% of the deposited sulfur is removed. Moreover, a stabilization of the remaining sulfates is observed. This result indicates that the sulfur poisoning is irreversible. In addition to a possible incomplete surface cleaning, the reducing treatment can induce some structural changes in the doped catalysts. Indeed, a slight deactivation of the PtBaMnCe/Al catalyst was also observed after the sulfur regeneration treatment without previous sulfating procedure.

Conclusion

In the first part of this study, it was shown that Ce, Mn addition to PtBa/Al led to an improvement of the NOx reduction (conversion and selectivity). A near total NOx conversion to N₂ can be observed at 400 °C with PtBaMnCe/Al. In the second part, the effect of sulfur in the gasoline on the catalyst performance was observed. The sulfur resistance of PtBaMnCe/Al catalyst is higher than reference PtBa/Al catalyst. Unfortunately, the NOx storage-reduction property cannot be fully recovered after the sulfur regeneration. These results clearly indicate that the lower sulfur level in fuels is indispensable for current and future emission control systems.

References
